Alkali metals and halogens

Which of the following statements is correct? (a) Hydrogen has same IP as alkali metals (b) Hydrogen has same electronegativity as halogens. (c) It has oxidation number of - 1 and + 1 ...

The elements of Group VIIA (new Group 17 - fluorine, chlorine, bromine, iodine, and astatine) are called the halogens (yellow column). The term "halogen" means "salt-former" because these elements will readily react with alkali metal and ...

Alkali metals are found in group 1 of the periodic table and have a +1 charge, while alkaline earth metals are found in group 2 and have a +2 charge. Alkali metals are highly ...

Reactivity: Group 1 metals, also known as alkali metals, are much more reactive than transition metals. They react vigorously with oxygen, water, and halogens, while transition ...

The noble gases are the elements that are nonreactive due to their complete valence electron shells, making them stable. In contrast, halogens, alkaline earth metals, and alkali metals are ...

Explanation Reactivity of Halogens The halogens, which are found in Group 17 of the periodic table, are known for their high reactivity. They love to react with alkali metals (Group 1), ...

Compare and contrast the characteristics of elements in Group 1 (Alkali metals) and Group 17 (Halogens) of the periodic table. l. Define the s, p, d, and f block elements and differentiate ...

Chemistry Questions Name the most electronegative and least electronegative element. What are amphoteric oxides? Give an example. Predict the formula of a stable binary compound of the ...

e. Hydrogen resembles both the transition metals and the halogens Correct. Hydrogen"s properties resembled both alkali metals and halogens, making its position in the periodic table ...

Halogen, any of the six nonmetallic elements that constitute Group 17 (Group VIIa) of the periodic table. The halogen elements are fluorine (F), chlorine (Cl), bromine (Br), iodine (I), astatine (At), and tennessine (Ts). Learn ...

Sodium is an alkali metal, and alkali metals are well-known for their ability to easily lose electrons. They have a single electron in their outermost shell, and losing this electron allows them to ...

Noble Gases in Group 18 have a full outer shell, making them very stable and unreactive. Unlike alkali metals and halogens, they do not readily participate in chemical reactions. Hence, Group ...

Reason 1: Position of Hydrogen: Hydrogen resembles both alkali metals (Group 1) and halogens (Group 17) in its properties. It was difficult to decide whether to place it with alkali metals or ...

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